



$$pH = -\lg\left(\frac{c(H_3O^+)}{\frac{mol}{L}}\right)$$

$$-pH = \lg\left(\frac{c(H_3O^+)}{mol/L}\right)$$

$$\lg^{-1}(-pH) = \frac{c(H_3O^+)}{mol/L}$$

$$\lg^{-1}(-pH) * mol/L = c(H_3O^+)$$

$$\lg^{-1}(-5,5) * mol/L = 0,0000032 mol/L$$